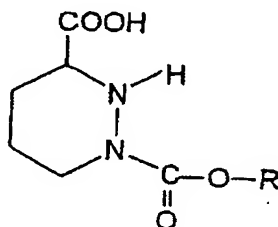


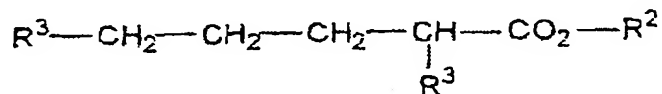
CLAIMS

1) A process for preparing the hexahydropyridazine-3-carboxylic acid derivatives of formula (I)

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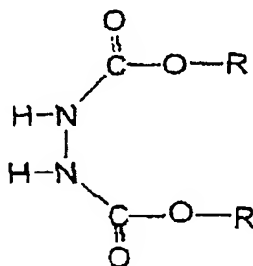


in which R represents a saturated or unsaturated, substituted or unsubstituted alkyl radical, a substituted or unsubstituted aralkyl radical or a substituted or unsubstituted aryl radical, characterized in that a compound of formula (II)



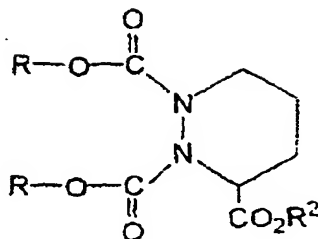
in which R<sup>2</sup> represents a substituted or unsubstituted alkyl radical, and R<sup>3</sup> represents a halogen atom or a nucleofugal organic group, is reacted with a compound of formula (III)

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in which R has the above meaning, in the presence of a base with a pK of greater than or equal to 8.5, in an organic solvent chosen from ketones, to obtain the tetrahydro-1,2,3-pyridazine-tricarboxylate intermediate compound of formula (IV)

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in which R and R<sup>2</sup> have the above meanings, which is not isolated and which is treated with a basic aqueous medium, to obtain the hexahydropyridazine-3-carboxylic acid derivative of formula (I).

2) The process as claimed in claim 1, wherein the organic solvent is chosen from acetone, methyl ethyl ketone, methyl isobutyl ketone, methyl tert-butyl ketone and diisopropyl ketone, and mixtures thereof.

3) The process as claimed in claim 1 or 2, characterized in that the base used in the first reaction is chosen from the group consisting of alkali metal carbonates and tertiary amines.

4) The process as claimed in any one of the preceding claims, characterized in that the solvent is acetone.

5) The process as claimed in any one of the preceding claims, characterized in that the base used in the first reaction is potassium carbonate.

6) The process as claimed in any one of the preceding claims, characterized in that the base used for the second reaction is chosen from alkali metal hydroxides and alkali metal or alkaline-earth metal alkoxides.

7) The process as claimed in the preceding claim, characterized in that the alkali metal hydroxides are used in aqueous solution.

8) The process as claimed in claim 6 or 7, characterized in that the mineral base is sodium hydroxide or potassium hydroxide.

5 9) The process as claimed in any one of the preceding claims, characterized in that, for the second reaction, the temperature is between 25°C and 55°C and the volume of water is between 1 and 10 liters per kilogram of compound of formula (III).

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10) The process as claimed in the preceding claim, characterized in that the reaction is performed by applying different successive temperature stages within the range.

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11) The process as claimed in any one of the preceding claims, characterized in that the compound of formula (I) is obtained in crystalline form by mixing the reaction medium with a solvent in which the compound of  
20 formula (I) is insoluble and which is a diluent for alcohols, and by bringing the pH of the medium to a value of between 0.5 and 2 using an acid.

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12) The process as claimed in the preceding claim, characterized in that the solvent is chosen from the group consisting of aromatic hydrocarbons, aliphatic hydrocarbons, ethers and acetates.

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13) The process as claimed in claim 11 or 12, characterized in that the acid is hydrochloric acid.

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14) The process as claimed in any one of the preceding claims, characterized in that R<sup>1</sup> represents the phenyl or naphthyl radical, R<sup>2</sup> represents a C<sub>1</sub> to C<sub>4</sub> alkyl radical and R<sup>3</sup> represents a halogen atom.

15) The process as claimed in the preceding claim, characterized in that R<sup>1</sup> represents the phenyl radical,

$R^2$  represents a methyl radical and  $R^3$  represents a bromine atom.